

# Reaction Monitoring using UV-Resonance Raman Spectroscopy

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## Application Note

### Introduction

Raman Spectroscopy is an optical technique that is especially suited to in-situ reaction monitoring in chemical synthesis. However, when low concentrations of the reactants are involved it is no longer feasible due to the intrinsic weakness of the Raman scattering of the reagents relative to the solvent. To circumvent this problem of sensitivity one can obtain an enhanced Raman signal of specific reaction components using techniques such as Surface Enhanced and Resonance Raman Spectroscopy. The use of UV-Resonance Raman is illustrated here by work carried by our group at the University of Groningen [1]. Their aim was to demonstrate the feasibility of using UV-RR as an online reaction monitoring technique for low temperature catalytic oxidation of organic substrates (e.g., bleaching) in aqueous environments. An important industrial application of the use of low temperature catalysts is the bleaching of cotton fabrics.

Abdolazadeh et al [1] used the epoxidation of alkenes, catalysed by a manganese complex, as a proof of concept model reaction. Manganese complexes of the ligand 1,4,7-trimethyl-1,4,7-triazacyclononane (TMTACN) were identified as highly effective catalysts for low temperature bleaching in the early nineties [2, 3]. However, due to the low substrate solubility in aqueous conditions one is generally dealing with low substrate concentrations (<10 mM). A key challenge in reaction monitoring in such conditions is having the 'sensitivity' to work with such low concentrations.

A schematic of the chemical reaction used for the demonstration is shown in figure 1. It illustrates the oxidation of 4-vinylbenzoic acid in water with  $H_2O_2$  catalysed by the manganese complex. The reaction progress was monitored by recording the resonance Raman spectrum and monitoring the change in intensity of selected spectral features relative to an internal standard ( $Na_2SO_4$ ).

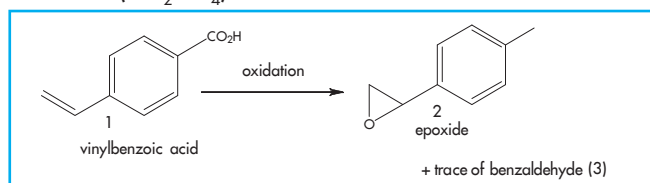


Figure 1: Schematic of the model reaction for demonstrating the technique. 4-Vinyl-benzoic acid is oxidised in water with  $H_2O_2$  to form the reaction products, an epoxide and traces of the corresponding carboxybenzaldehyde. A manganese complex was used as the catalyst.

### Experimental Setup

The main components used in the experimental setup are illustrated in figure 2. The sample solution of 1 mM 4-vinyl-benzoic acid dissolved in water upon addition of  $H_2O_2$  and an internal standard (0.05 M  $Na_2SO_4$ ) was held in a quartz cuvette (1 cm pathlength 3 mL volume) and at time zero the manganese catalyst (final concentration 0.01 mM) was added. The reaction was carried out at 18 °C. An excitation wavelength of 266 nm was chosen as it coincides with the strong UV absorption bands of the 4-vinyl-benzoic acid and (to a lesser extent) the epoxide and aldehyde products (see figure 4), which facilitates resonance enhancement of the Raman scattering from the reactant and products. 266 nm laser light, the fourth harmonic of an Innolas Spitlight200 Nd:Yag ns-pulsed laser (10 Hz, 4-5 ns pulse width), was delivered using dichroic mirrors (to remove residual 532 nm light) and quartz prisms, providing an average power at sample of 1-2 mW. A 180° backscattering geometry was employed with quartz planoconfocal lenses being used for focusing and collimating (see optical table in figure 2). The collected scattering was focused into a spectrograph (Shamrock SR-303i-B, 300 nm blaze 2400 l/mm grating) and onto a UV-sensitive iDus CCD detector (DV420A-BU2). The BU2 sensor is a back-thinned sensor with AR coating optimised for the deeper UV circa 250 nm, making it ideal for this spectral region. A long-pass filter (266 nm RazorEdge® ultrasteep long-pass edge filter from Semrock) was placed just in front of the spectrograph entrance slit to prevent the Rayleigh scattered light from entering the spectrograph. Raman spectra with typical acquisitions settings of 60 accumulations of 2 s exposure were captured at intervals of 15 minutes.

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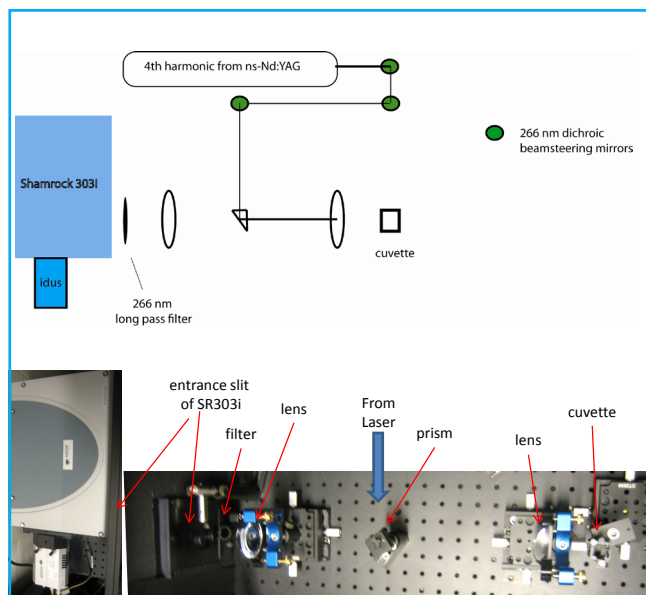


Figure 2: Schematic of the experimental setup. A picture of the optical layout is included at the bottom.

### Results

Spectra taken on pure samples of 4-vinyl-benzoic acid and the epoxide and aldehyde products – see figures 3 and 4 – show the main Raman feature used to monitor the reaction progress at a Raman shift of  $1631\text{ cm}^{-1}$ . Solid state samples excited at  $785\text{ nm}$  produced the spectra in figure 3 whilst solutions ( $1\text{ mM}$ ) were used for those of figure 4A. The corresponding UV absorption spectra are shown in figure 4B.

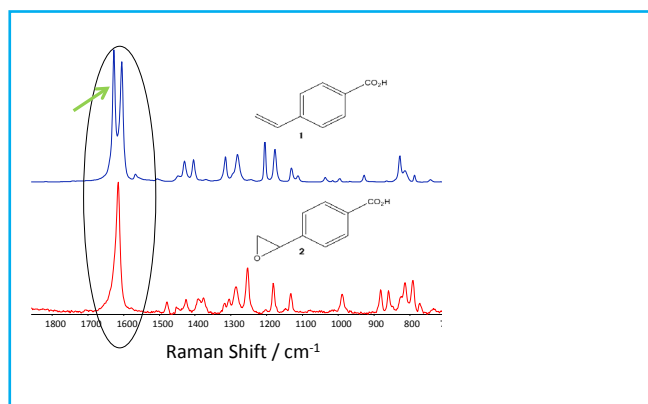


Figure 3: Comparison of the Raman 'fingerprints' for the 4-vinyl-benzoic acid and the epoxide product. Spectra were taken on solid state samples using an excitation wavelength of  $785\text{ nm}$ . The key feature used for monitoring the reaction kinetics is highlighted at a Raman band around  $\sim 1631\text{ cm}^{-1}$ .

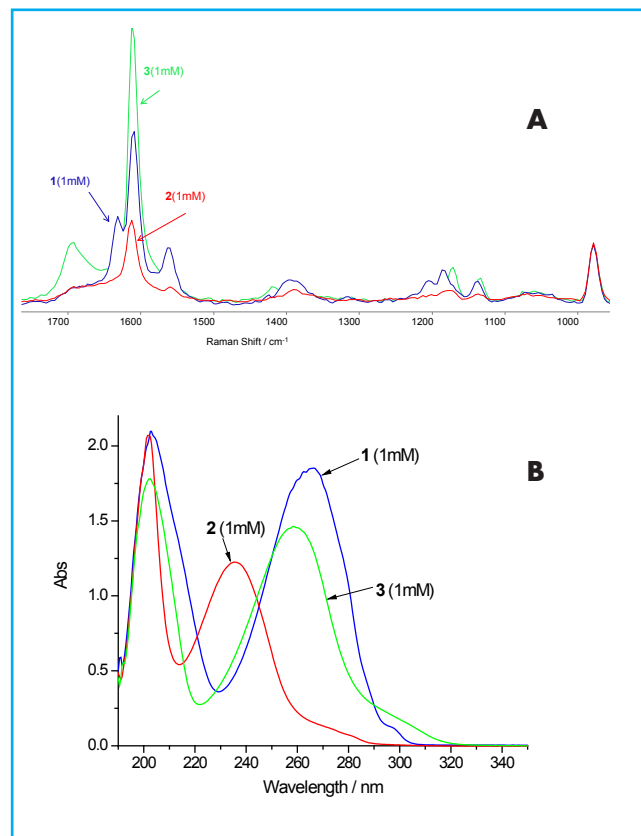


Figure 4: A shows the resonance Raman spectra (in  $0.1\text{ M NaHCO}_3$  buffer,  $\text{pH}=8$ ) of pure samples of 4-vinyl-benzoic acid (**1**,  $1\text{ mM}$ ), the epoxide product (**2**,  $1\text{ mM}$ ) and the carboxy-benzaldehyde (**3**,  $1\text{ mM}$ ) product (internal reference  $\text{Na}_2\text{SO}_4$   $0.25\text{ M}$ ). The corresponding UV/Vis absorption spectra are shown in B.

A series of spectra taken at different times under the real experimental conditions, with concentrations of  $1\text{ mM}$ , are shown in figure 5. The intensity of the Raman band at  $1631\text{ cm}^{-1}$  can be seen to fall as the reaction progresses. Figure 6 shows the variation of the intensity as a function of time.

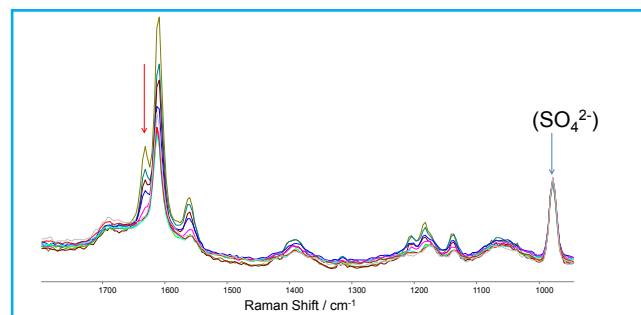


Figure 5: A series of captured Raman spectra illustrating how the variation in intensity of the feature at  $1631\text{ cm}^{-1}$  facilitates monitoring of the reaction process through time. The standard Raman band from  $\text{Na}_2\text{SO}_4$  ( $0.05\text{ M}$ ) is evident on the right.



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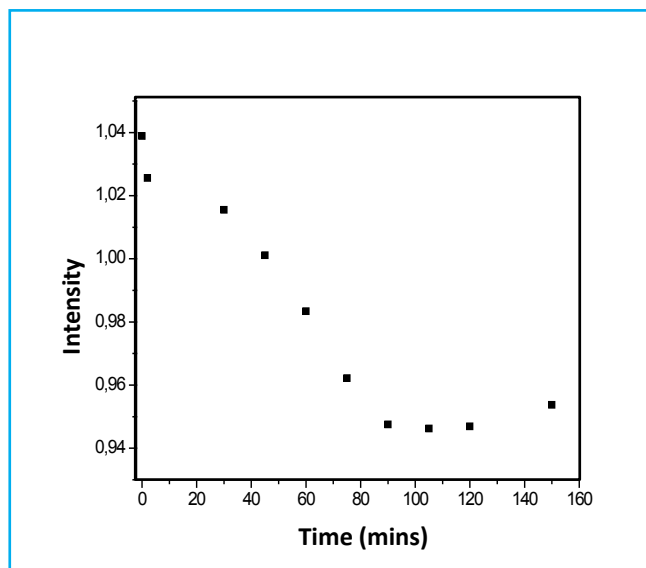


Figure 6: Variation with time of the intensity of the Raman band at 1631 cm<sup>-1</sup>.

The results in figure 6 demonstrate the potential of UV-Resonance Raman Spectroscopy as a real-time on-line reaction monitoring technique [4], particularly when working with low concentrations of reactants (<1 mM). Among the key enablers of such screening techniques for low concentration species will be the optimisation and enhancement of the signal - as with Resonance Raman in this case, efficient light collection and sensitive multichannel detectors in the UV - as with the BU2 sensor.

### References

- [1] S. Abdolazadeh, A. Draksharapu, J. W. de Boer, R. Hage and W. R. Browne. 'Reaction monitoring with UV-Raman spectroscopy of the oxidation of alkenes with H<sub>2</sub>O<sub>2</sub> in water catalysed by manganese complexes', Poster presentation–XIX ISPPCC, Strasbourg June (2011)
- [2] R. Hage, J. E. Wilburg, J. Kerschner, J. H. Koek, E. L. M. Lempers, R. J. Martens, U. S. Racherla, S. W. Russell, T. Swarthoff, M. R. P. van Vliet, J. B. Warnaar, L. van der Wolf, B. Krijnen, *Nature* 369 (1994) 637-639.
- [3] R. Hage, A. Lienke, *Angew. Chem. Int. Ed.* 45, (2006) 206-222.

For further information – see web links

<http://www.stw.nl/Projecten/G/goc/11059.htm> and  
<http://www.rug.nl/scheikunde/onderzoek/scholen/stratingh/groepen/Browne/object20627553>

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